Chemistry 141 Name

Dr. Cary Willard

Double Quiz A – Quantum (42 points) November 2, 2010

1. (15 points) The threshold frequency of titanium is 1.05 x 1015/sec. Answer the following questions regarding titanium
   1. What is the minimum wavelength required to ionize an atom of titanium?
   2. What is the minimum energy required to ionize an atom of titanium?
   3. How many photons would be required to ionize 4.25 grams of titanium?
   4. What is the ionization energy in kJ/mol for titanium?
   5. If a sample of titanium was exposed to light with a wavelength of 215 nm, calculate the kinetic energy of the electrons emitted.
2. (3 points) Ultraviolet radiation causes skin damage that may lead to cancer, but exposure to infrared radiation does not seem to cause skin cancer. Why do you think this is so?
3. (3 points) Explain the difference between a ground-state H atom and an excited-state H atom.
4. (9 points) The shorthand configuration for an atom of tellurium is [Kr] 5s2 4d10 5p4?
   1. How many valence electrons are there in an atom of Te?
   2. Draw an orbital diagram for an atom of Te showing all of the electrons beyond Kr. (You know the one, shows all of the little boxes representing orbitals with arrows representing the electrons.)

Give possible quantum numbers for the 5p electrons in Te.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
|  | n | l | ml | ms |
| Electron 1 |  |  |  |  |
| Electron 2 |  |  |  |  |
| Electron 3 |  |  |  |  |
| Electron 4 |  |  |  |  |

1. (3 points) How many quantum numbers are needed to identify an electron in an atom?
2. (3 points) Explain why so many transition metals form ions with a +2 charge.
3. (3 points) Write the complete electron configuration for an atom of P.
4. (3 points) Write the shorthand electronic configuration for an atom of europium as predicted from the periodic table.

Formulas

|  |  |  |
| --- | --- | --- |
| Kinetic energy = ½ mv2  w = -PΔV  ΔG = ΔH - TΔS  w=dxF  C = q/ΔT  ΔGo = -nFEo  ΔG = - RTlnK  E = mc2  E = IR | PV = nRT  Ptotal = P1+P2+P3+…  P1=X1\*Ptotal  Ba(Na)2 = fruit  Ptotal = P1 + P2 + P3 + …  M = mol/L  m = mol/kg solvent  Xi = moli/ moltotal  u = (3RT/MW)½  Rate ∝ (MW)-½ | HΨ=EΨ  Amp = C/sec  π= iMRT  E = hν = hc/λ  E = nhν = nhc/λ  M1V1 = M2V2  Psoln = (Psolv)(Xsolv)  ΔTf = i(kf)(m)  ΔTb = i(kb)(m) |



Constants

|  |  |
| --- | --- |
| F = 9.65 x 104 C  h = 6.626 x 10-34 J sec  c= 2.9979 x 108 m/sec  e = 1.602 x 10-19 C  NA = 6.022 x 1023/mol  k = 1.381 x 10-23 J/K  1 kcal = 4.184 kJ  K = oC + 273.16  Kw = 1.0 x 10-14M2 | mass electron = 9.109 x 10-31 kg  Standard Temperature and Pressure = 0oC and 1 atm  R = 0.0821 L atm/mol K= 8.314 J/K mol= 1.987 cal.mol K = 62.4 L torr/mol K  760 torr = 760 mm Hg = 1.00 atm = 101 kPa = 14.6 psi = 30 in Hg |

Chemistry 141 Name

Dr. Cary Willard

Double Quiz B – Quantum (42 points) November 2, 2010

1. (15 points) The threshold frequency of silicon is 1.97 x 1015/sec. Answer the following questions regarding silicon
2. What is the minimum wavelength required to ionize an atom of silicon?
3. What is the minimum energy required to ionize an atom of silicon?
4. How many photons would be required to ionize 4.25 grams of silicon?
5. What is the ionization energy in kJ/mol forsilicon?
6. If a sample of silicon was exposed to light with a wavelength of 115 nm, calculate the kinetic energy of the electrons emitted.
7. (3 points) Ultraviolet radiation causes skin damage that may lead to cancer, but exposure to infrared radiation does not seem to cause skin cancer. Why do you think this is so?
8. (3 points) Explain the difference between a ground-state H atom and an excited-state H atom.
9. (9 points) The shorthand configuration for an atom of tellurium is [Kr] 5s2 4d10 5p4?
10. How many valence electrons are there in an atom of Te?
11. Draw an orbital diagram for an atom of Te showing all of the electrons beyond Kr. (You know the one, shows all of the little boxes representing orbitals with arrows representing the electrons.)

Give possible quantum numbers for the 5p electrons in Te.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
|  | n | l | ml | ms |
| Electron 1 |  |  |  |  |
| Electron 2 |  |  |  |  |
| Electron 3 |  |  |  |  |
| Electron 4 |  |  |  |  |

1. (3 points) How many quantum numbers are needed to identify an electron in an atom?
2. (3 points) Explain why so many transition metals form ions with a +2 charge.
3. (3 points) Write the complete electron configuration for an atom of Al.
4. (3 points) Write the shorthand electronic configuration for an atom of gadolinium as predicted from the periodic table.

Formulas

|  |  |  |
| --- | --- | --- |
| Kinetic energy = ½ mv2  w = -PΔV  ΔG = ΔH - TΔS  w=dxF  C = q/ΔT  ΔGo = -nFEo  ΔG = - RTlnK  E = mc2  E = IR | PV = nRT  Ptotal = P1+P2+P3+…  P1=X1\*Ptotal  Ba(Na)2 = fruit  Ptotal = P1 + P2 + P3 + …  M = mol/L  m = mol/kg solvent  Xi = moli/ moltotal  u = (3RT/MW)½  Rate ∝ (MW)-½ | HΨ=EΨ  Amp = C/sec  π= iMRT  E = hν = hc/λ  E = nhν = nhc/λ  M1V1 = M2V2  Psoln = (Psolv)(Xsolv)  ΔTf = i(kf)(m)  ΔTb = i(kb)(m) |



Constants

|  |  |
| --- | --- |
| F = 9.65 x 104 C  h = 6.626 x 10-34 J sec  c= 2.9979 x 108 m/sec  e = 1.602 x 10-19 C  NA = 6.022 x 1023/mol  k = 1.381 x 10-23 J/K  1 kcal = 4.184 kJ  K = oC + 273.16  Kw = 1.0 x 10-14M2 | mass electron = 9.109 x 10-31 kg  Standard Temperature and Pressure = 0oC and 1 atm  R = 0.0821 L atm/mol K= 8.314 J/K mol= 1.987 cal.mol K = 62.4 L torr/mol K  760 torr = 760 mm Hg = 1.00 atm = 101 kPa = 14.6 psi = 30 in Hg |